

Revision Notes

Chapter – 6

Physical and Chemical Changes

- Changes can be of two types, physical and chemical.
1. **Physical changes** are changes in the physical properties of substances. No new substances are formed in these changes. These changes may be reversible.
Example: crushing a can, glowing of an electric bulb, tearing of paper, mixing of sand and water.
 2. **Chemical Changes** are changes in which the composition and chemical properties of the substance get changed. In chemical changes new substances are produced. This change is permanent and irreversible.
Example: burning of a candle, formation of curd from milk, ripening of fruits.
- **Chemical Reactions in daily life:**
1. **Rusting of Iron:** Rusting is the process in which iron turns into iron oxide. It happens when iron comes into contact with water and oxygen. The process is a type of corrosion that occurs easily under natural conditions.

Prevention of Rusting:

- (i) By Painting
 - (ii) By Oiling and greasing
 - (iii) By Chromium plating
 - (iv) By Galvanizing
 - (v) By Alloying
2. **Cooking of food:** Cooking causes breakdown of complex molecules of carbohydrates, fats and proteins into smaller molecules.
It is regarded as a decomposition reaction. Cooked food is easier to digest than uncooked food.

3. **Decay of Organic Substances:** Microorganisms like fungi and bacteria produce enzymes which break down complex organic compounds into smaller substances.

It is also regarded as a decomposition reaction.

- Some substances can be obtained in pure state from their solutions by crystallization.

